Practice:

1. In nature, silicon is composed of three isotopes. These isotopes are 92.23% Si-28, 4.67% Si-29 and 3.10% Si-30. Calculate the average atomic mass of silicon.

2. The two stable isotopes of boron exist in the following proportions: 19.78% $^{10}_{5}B$ (10.01 u) and 80.22% $^{11}_{5}B$ (11.01 u). Calculate the average atomic mass of boron.

3. Copper exists at two naturally occurring isotopes: $^{63}_{29}Cu$ (62.93 u) and $^{65}_{29}Cu$ (64.93 u). These isotopes have isotopic abundances of 69.1% and 30.9% respectively. Calculate the average atomic mass of copper.

4. Lead occurs naturally as four isotopes. These isotopes are 1.37% $^{204}_{82}Pb$ (204.0 u), 26.26% $^{206}_{82}Pb$ (206.0 u), 20.82% $^{207}_{82}Pb$ (207.0 u), and 51.55% $^{208}_{82}Pb$ (208.0 u). Calculate the average atomic mass of lead.